

STABILIZATION OF POLYMERS
AFTER EXPOSURE TO OXIDATION

Background of the Invention

5 This application is a continuation-in-part of application Serial Number
09/272,843, filed March 19, 1999, now abandoned.

10 This invention relates to the use of stabilizers to reduce the discoloration of
certain polymers after they have been exposed to oxidation. In particular, it relates
to the use of derivatives of certain unsaturated diols for that purpose.

15 A number of different polymers, including polyvinyl chloride (PVC),
polycarbonates, polyurethane, polyethylene, and polypropylene, are used to make
various medical devices and as packaging for food. Sterilization can be
accomplished by exposing the devices or packages to gamma radiation. However,
the gamma radiation can degrade or yellow the polymer, making it unsuitable or less
acceptable for certain applications. Polymers such as polyethylene, polypropylene,
and PVC can also degrade and yellow after heating.

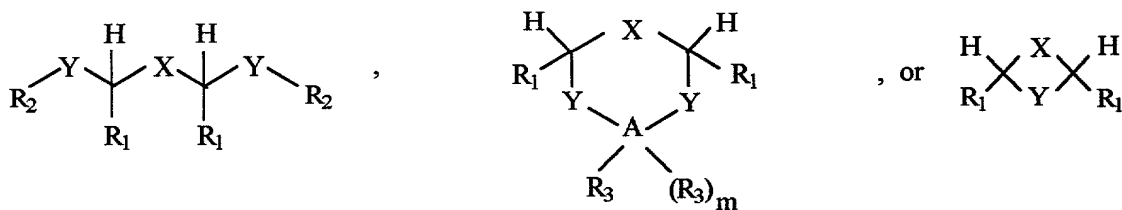
Summary of the Invention

We have discovered that when a particular type of stabilizer is added to
certain polymers, the polymers degrade and yellow substantially less after exposure
to oxidation. Some of the stabilizers of this invention are novel compounds while

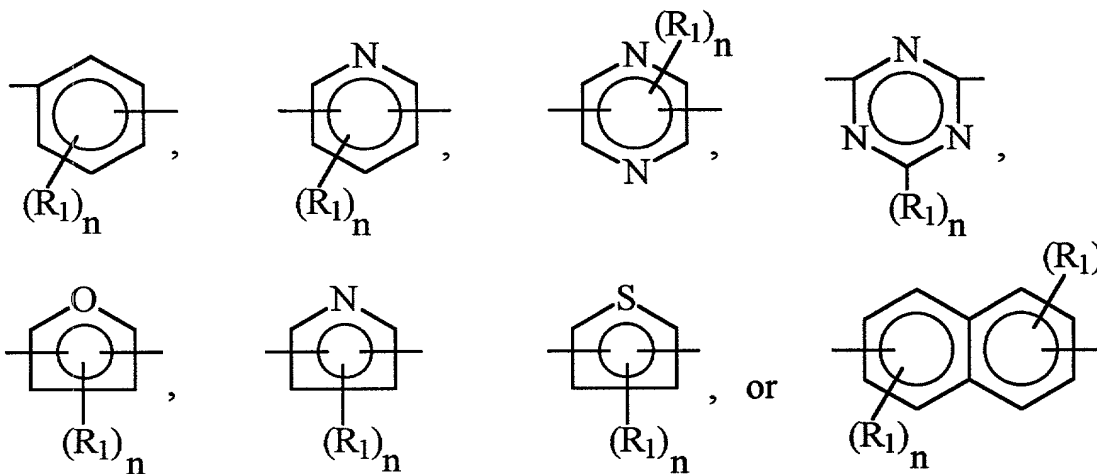
others are commercially available. Some of these stabilizers are polymeric and resist leaching or migration from the polymer, thereby further extending the life of the polymer.

Description of the Preferred Embodiments

5 The stabilizers of this invention have the general formula:



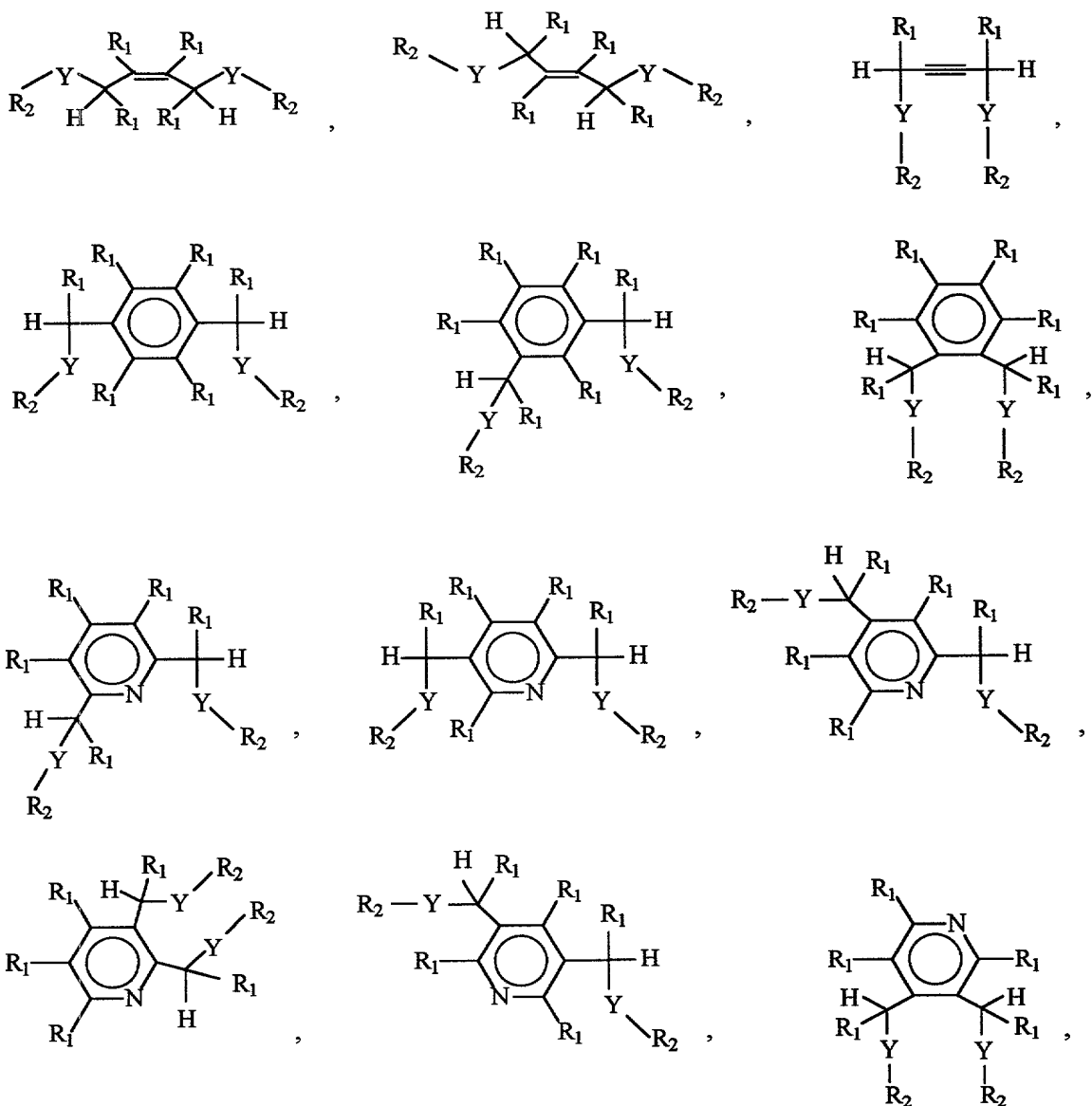
where A is C, P, Sn, Si, or B, X is $-\text{R}_1\text{C}=\text{CR}_1-$, $-\text{C}\equiv\text{C}-$,



each Y is independently selected from O and S; each R is independently selected from hydrogen, alkyl from C_1 to C_{20} , aryl from C_6 to C_{20} , alkaryl from C_7 to C_{20} , and

aralkyl from C_7 to C_{20} ; each R_1 is independently selected from R, OR, RCO, ROCO, $ROCO_2$, $P(R)_2$, $P(OR)_2$, $PR(OR)$, $N(R)_2$, $(R)_2NCO$, $(R)_2NCO_2$, SR, and halogen, where two R_1 groups can be bridged together to form a ring; each R_2 is independently selected from R, RCO, ROCO, $P(OR)_2$, $Sn(R)_p(OR)_{3-p}$, $Sn(R)_p(OCOR)_{3-p}$, $Si(R)_p(OR)_{3-p}$, and $B(R)_p(OR)_{2-p}$, except that when two Y's are O and X is $-R_1C=CR_1-$, at least one R_2 is not hydrogen, each R_3 is independently selected from R, RCO, ROCO, $ROCO_2$, OR, SR, $N(R)_2$, $OP(R)_2$, and $OP(OR)_2$; m is 0 when A is P or B and is 1 when A is Sn, Si, or C; n is 0 to 4, depending on the number of available sites; and p is 0 to 3 for the tin stabilizers and 0 to 2 for the boron stabilizers. Two R_1 or two R_2 groups or an R_1 group and an R_2 group can be bridged together to form a ring. Groups that can bridge R_1 or R_2 groups together to form a ring include R_4 , $-HC=CH-$, $-CH_2O-$, $-CH_2NH-$, $-HC=N-$, and $-CH_2S$, where R_4 is alkylene from C_1 to C_{20} , arylene from C_6 to C_{20} , (aryl)alkylene from C_7 to C_{20} , (alkyl)arylene from C_7 to C_{20} , alkanediyl from C_1 to C_{20} , (aryl)alkanediyl from C_7 to C_{20} , $-CO-(alkylene)-CO-$ from C_1 to C_{20} , $-CO-arylene-CO-$ from C_6 to C_{20} , $-CO-(aryl)alkylene-CO-$ from C_7 to C_{20} , $-CO-(alkyl)arylene-CO-$ from C_7 to C_{20} , $Si(R)_2$, $SiR(OR)$, $Si(OR)_2$, $P(OR)$, $B(OR)$, $Sn(R)_2$, $SnR(OR)$, or $SnR(O-CO-R)$. Preferably, A is C; X is $-HC=CH-$; Y is O; R is hydrogen, alkyl from C_1 to C_{12} , aryl from C_6 to C_{12} , alkaryl from C_7 to C_{12} , or aralkyl from C_7 to C_{12} , and is most preferably benzyl; R_1 is H; R_2 is R; R_3 is R; R_4 is alkylene from C_1 to C_8 , (aryl)alkylene from C_7 to C_8 , or $-CO-(aryl)alkylene-CO-$ from C_7 to C_8 ; and p is 0 because those compounds work well and are readily available. Stabilizers where each R_1 is hydrogen are especially preferred as they are more effective.

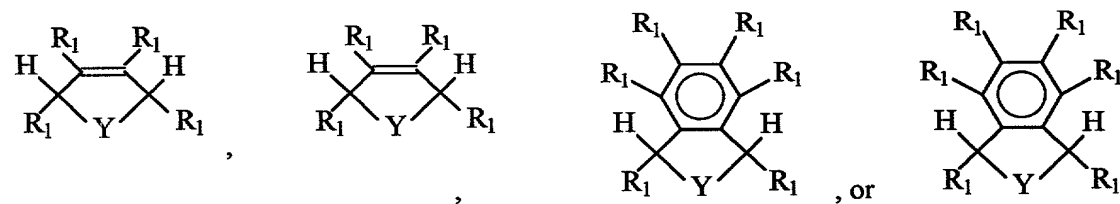
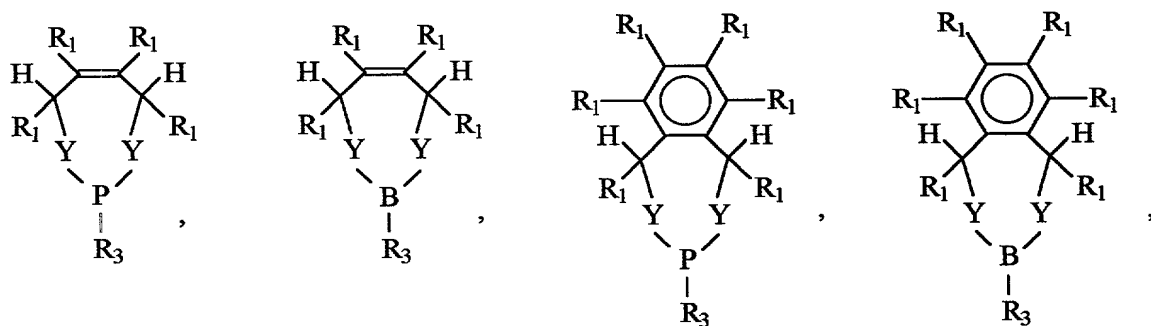
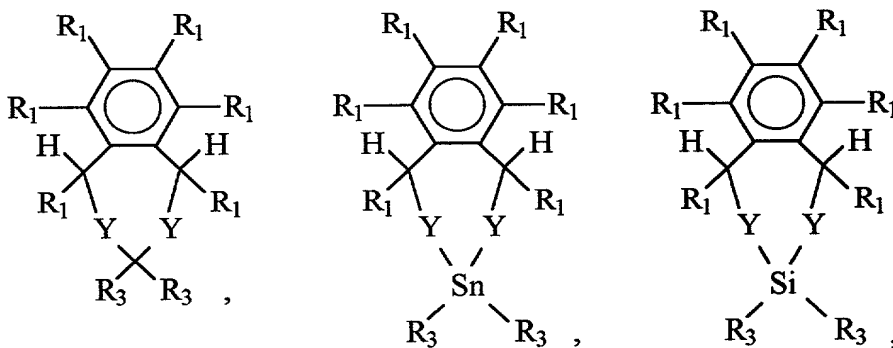
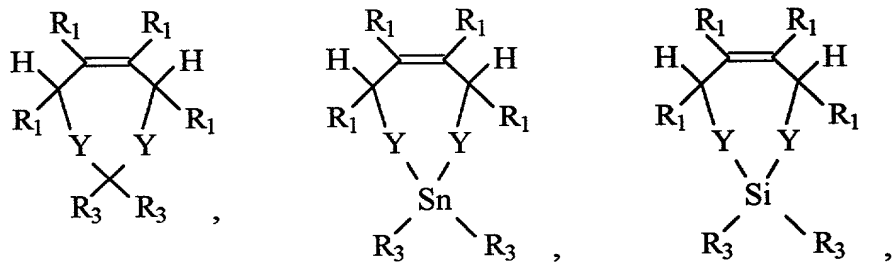
When X is $-R_1C=CR_1-$ the stabilizers can be either cis or trans, but the cis stabilizers are preferred because they are less expensive. The following are examples of some of the stabilizers included within the scope of the above formulas:



Chemical structures 1-6 and 7 are defined as follows:

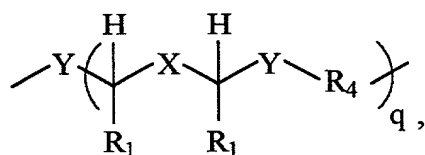
- Structure 1:** A benzene ring with substituents R_1 at positions 1, 3, and 5, and $Y-R_2$ at positions 2 and 4. Hydrogen atoms are at positions 6 and the other two positions.
- Structure 2:** A benzene ring with substituents R_1 at positions 1, 3, and 5, and $Y-R_2$ at positions 2 and 4. Hydrogen atoms are at positions 6 and the other two positions.
- Structure 3:** A pyridine ring with substituents R_1 at positions 2, 4, and 6, and $Y-R_2$ at positions 3 and 5. Hydrogen atoms are at positions 1 and the other two positions.
- Structure 4:** A benzene ring with substituents R_1 at positions 1, 3, and 5, and $Y-R_2$ at positions 2 and 4. Hydrogen atoms are at positions 6 and the other two positions.
- Structure 5:** A benzene ring with substituents R_1 at positions 1, 3, and 5, and $Y-R_2$ at positions 2 and 4. Hydrogen atoms are at positions 6 and the other two positions.
- Structure 6:** A benzene ring with substituents R_1 at positions 1, 3, and 5, and $Y-R_2$ at positions 2 and 4. Hydrogen atoms are at positions 6 and the other two positions.
- Structure 7:** A naphthalene ring with substituents R_1 at positions 1, 3, and 5, and $Y-R_2$ at positions 2 and 4. Hydrogen atoms are at positions 6 and the other two positions.

Examples of stabilizers having a cyclic structure include:

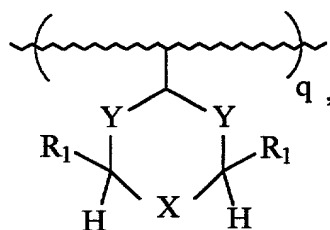


Examples of stabilizers having the above structure include 4,7-dihydro-1,3-dioxepins, ethers and esters of butene-1,4-diol, and phthalans. These stabilizers can be prepared by condensing an allylic diol, or an aromatic analog of an allylic diol, with an aldehyde, ketone, acid, acid halide, ester, alkyl halide, or alcohol or by other reactions known to those skilled in the art, to form derivatives. See, for example, "A Stereospecific Route to Aziridinomitosanes: The Synthesis of Novel Mitomycin Congeners," by Samuel Danishefsky et al. *J. Am. Chem. Soc.* **1985**, *107*, 3891-3898.

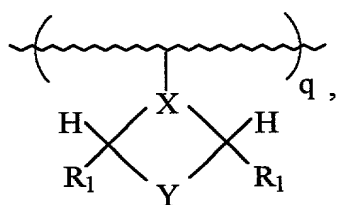
The stabilizers can be oligomeric or polymeric versions of the compounds shown in the first formula that have the structure:



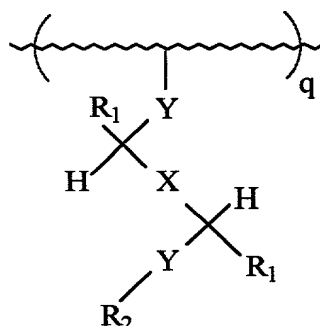
A



B



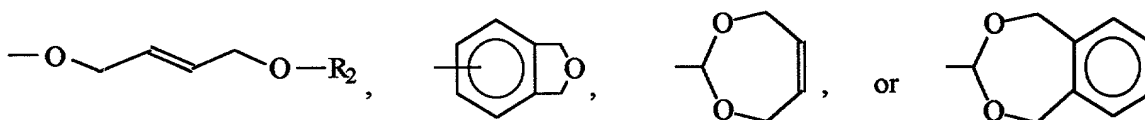
or



C

D

In the above formulas, q can be 1 to 1000, but is preferably 1 to 5. Examples of polymers having formula A include poly(2-butenylene adipate), poly(2-butenylene-terephthalate, and poly[oxy(2-butenylene)]. In formulas B, C, and D, the preferred backbones are polyesters, polyethers, and polyketones, and the preferred pendant groups are



Examples of such compounds include the condensate of poly(ethylene-co-carbon monoxide) with 2-butene-1,4-diol. The polymeric stabilizers are expected to be more effective as they have less tendency to migrate or leach out of the stabilized polymer. Preparation of these polymers can be achieved by reactions known to those skilled in the art. For example, the polymeric esters can be prepared by reacting the diol (see the first formulas where Y is oxygen and R₁ is hydrogen) with a diester, diacid, diacid chloride, or dianhydride.

The stabilizers of this invention are effective against oxidation of PVC, polycarbonates, polyurethane, polypropylene, polyethylene, polyvinylidene chloride, polyamides, polyimides, polyethers, polyesters, polysiloxanes, polyurethanes, polysulfones, and polysulfides. The preferred polymers are PVC, polycarbonates, polypropylene, and polyethylene because those polymers are more frequently used in medical applications where they are subjected to gamma radiation; particularly preferred is PVC.

The polymer can be stabilized by the addition of about 0.005 to about 10 phr (parts by weight per 100 part by weight of the polymer) of the stabilizer to the polymer. Less stabilizer is less effective and more stabilizer offers little additional benefit. The preferred amount of stabilizer is about 0.2 to about 6 phr. The stabilizer can be added to a polymer in a variety of ways, such as mixing into the reactants at

the beginning or during polymerization. The stabilizer is preferably added after at least 70 wt% of the monomer has polymerized. The stabilizer can be added as a solid or with a solvent as a slurry or a solution. Common organic solvents such as N-methylpyrrolidone, diglyme, acetamide, acetone, methanol, ethanol, isopropanol, dimethylsulfoxide, and dimethylformamide can be used; water can also be used. Water miscible solvents, such as acetone, tetrahydrofuran, and methanol, are preferred for PVC. It is preferable to add the stabilizer in a solvent as that achieves a more uniform distribution of the stabilizer in the polymer. The stabilizer can also be added during the drying or compounding of the polymer. Various methods can be used for compounding, including milling, dry mixing, and extrusion. The stabilizers function as antioxidants to inhibit various forms of oxidation.

The following examples further illustrate this invention:

EXAMPLES 1 to 13

To a mixture of 150.00 g PVC (sold by Occidental Chemical Corporation as "Oxy 240"), 0.30 g stearic acid (used as a lubricant), 0.23 g of zinc and calcium mixed salts of mixed fatty acids (used as a heat stabilizer; sold by Witco as "Mark 152 S"), 97.50 g dioctyl phthalate (used as a heat stabilizer and to increase flexibility), 15.00 g epoxidized soy bean oil (used as an HCl scavenger to reduce degradation; sold by Witco as "Drapex 6.8"), was added a stabilizer. The mixture was thoroughly blended and hot milled at 300°F (149°C) for 5 minutes. The resulting PVC sheet was cut and pressed into a 4"x3"x1/4" (10x8x0.6 cm) plaque at 330°F

(166°C). The plaque was divided into two smaller pieces. One was saved for comparison and one was subjected to γ radiation at a dose of 50 kGy. The irradiated piece was again divided into two pieces and one of these pieces was oven aged at 50°C for 48 hours. All of the samples were measured for yellowness index with a Macbeth 2020 Plus Color Eye Spectrometer, as described by the Hunter equations (see "The Measurement of Appearance" by Richard S. Hunter, John Wiley & Sons, New York, 1975). The stabilizers used in Examples 1 to 6 and 8 to 11 were purchased and the stabilizer used in Example 7 was made according to U.S. Patent No. 3,280,148. The following table gives the stabilizers used and the results.

Example	Stabilizer	Amount (g)	Yellowness		
			Initial	After Radiation	After Aging
1	2,5-dimethoxy-2,5-dihydrofuran	0.64	17.1	47.8	69.3
Control	none		18.4	52.8	74.1
2	phthalan	1.34	17.5	31.9	49.5
Control	none		18.1	50.2	75.5
3*	cis-4-benzyloxy-2-buten-1-ol	1.67	16.6	21.1	30.7
Control	none		16.3	43.3	66.6
4	1,4-dibenzyloxy-2-butene	2.23	17.6	23.9	38.2
Control	none		16.3	43.3	66.6
5	trans-2,3-dibromo-2-butene-1,4-diol	2.06	31.8	37.6	61.7
Control	none		16.3	43.3	66.6
6	4,7-dihydro-2-phenyl-1,3-dioxepin	1.47	16.1	26.0	41.1
Control	none		16.3	43.3	66.6
7	2,2'-(1,4-phenylene)bis(4,7-dihydro-1,3-dioxepin)	2.40	24.3	36.5	40.9
Control	none		18.1	50.2	75.5
8	1,5-dihydro-3-methoxy-2,4-benzodioxepin	2.03	16.7	36.0	49.8
Control	none		18.1	50.2	75.5
9	2-butyne-1,4-diol	0.96	18.0	39.9	62.9
Control	none		18.1	50.2	75.5
10	1,4-bis(2-hydroxyethoxy)-2-butyne	1.99	32.0	31.6	40.2
Control	none		18.1	50.2	75.5
11	1,4-benzenedimethanol	1.50	21.0	41.1	47.7
Control	none		18.1	50.2	75.5

* See also "Convenient One-pot Synthesis of (Z)-4-Benzyloxy-2-buten-1-ol," by S. V. Hiremath et al. *Indian Journal of Chemistry*, Vol. 27B, June 1988, p. 588.

As can be observed from the above table, after γ radiation the polymers that contained a stabilizer had significantly less discoloration than the control samples.